

PERCENTAGE COMPOSITION (MASS)

Question One: Calculate the percentage composition of the molecule 1-aminobutane. 1-aminobutane has the formula $C_4H_{11}N$.
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(N) = 14.0 \text{ g mol}^{-1}$

Question Two: Calculate the percentage composition of the malonic acid molecule. Malonic acid has the formula $C_3H_4O_4$
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(O) = 16.0 \text{ g mol}^{-1}$

Question Three: Calculate the percentage composition of the sebacic acid molecule. Sebacic acid has the formula $C_{10}H_{18}O_4$
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(O) = 16.0 \text{ g mol}^{-1}$

Question Four: Calculate the percentage composition of the putrescine molecule. Putrescine has the formula $C_4H_{12}N_2$
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(N) = 14.0 \text{ g mol}^{-1}$

Question Five: Calculate the percentage composition of the compound Na_2CO_3
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(Na) = 23.0 \text{ g mol}^{-1}$ and $M(O) = 16.0 \text{ g mol}^{-1}$

Question Six: Calculate the percentage composition of the compound propanone. Propanone has the formula C_3H_6O .
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(O) = 16.0 \text{ g mol}^{-1}$

Question Seven: Calculate the percentage composition of 3-chloropropene. 3-chloropropene has the formula C_3H_5Cl
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(Cl) = 35.5 \text{ g mol}^{-1}$

Question Eight: Calculate the percentage composition of the aspirin molecule. Aspirin has the formula $C_9H_8O_4$
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.00 \text{ g mol}^{-1}$ and $M(O) = 16.0 \text{ g mol}^{-1}$

Question Nine: Calculate the percentage composition of benzyl chloride. Benzyl chloride has the formula C_7H_7Cl $M(C) = 12.0 \text{ g mol}^{-1}$,
 $M(H) = 1.00 \text{ g mol}^{-1}$, $M(Cl) = 35.5 \text{ g mol}^{-1}$

Question Ten: Calculate the percentage composition of paracetamol. Paracetamol has the formula $C_8H_9O_2N$
 $M(C) = 12.0 \text{ g mol}^{-1}$, $M(H) = 1.0 \text{ g mol}^{-1}$, $M(O) = 16.0 \text{ g mol}^{-1}$,
 $M(N) = 14.0 \text{ g mol}^{-1}$

ANSWERS

Q1	% C = 65.8%	% H = 15.1%	% N = 19.2%
Q2	% C = 34.6%	% H = 3.85%	% O = 61.5%
Q3	% C = 59.4%	% H = 8.91%	% O = 31.7%
Q4	% C = 54.5%	% H = 13.6%	% N = 31.8%
Q5	% C = 11.3%	% Na = 43.4%	% O = 45.3%
Q6	% C = 62.1%	% H = 10.3%	% O = 27.6%
Q7	% C = 47.1%	% H = 6.54%	% Cl = 46.4%
Q8	% C = 60.0%	% H = 4.44%	% O = 35.6%
Q9	% C = 66.4%	% H = 5.53%	% O = 28.1%
Q10	% C = 63.6%	% H = 5.96%	% O = 21.2% % N = 9.27%